## FORMATION OF $\beta$ -METHYLDEHYDROALANINE IN THE MILD ALKALINE TREATMENT OF THE ANTIBIOTIC PEPTIDOCYCLOLACTONE A-128-OP

I. G. Smirnova, A. B. Silaev, and G. S. Katrukha

UDC 615.779.931

The antibiotic A-128-OP is a tripeptide-octacyclolactone, the lactone bond in which is formed through the carboxy group of L-cis-3-hydroxyproline and the hydroxy group of L-threonine [1-3]. In addi-

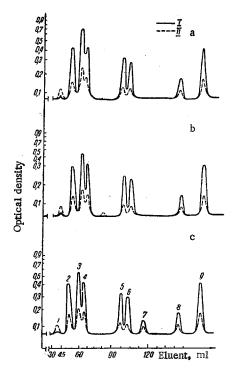


Fig. 1. Amino-acid composition of acid hydrolysates at optical densities of 570 (I) and 440 nm (II): a) antibiotic A-128-OP; b) antibiotic A-128-OP treated with 0.1 N NaOH; c) antibiotic A-128-OP after treatment with 0.1 N NaOH and hydrogenation over Pd/C. 1) trans-3-Hydroxyproline; 2) aspartic acid; 3) threonine, allothreonine, and cis-3-hydroxyproline; 4) serine; 5) glycine; 6) alanine; 7)  $\alpha$ aminobutyric acid; 8) erythro- $\beta$ -hydroxyleucine; 9) norleucine (standard). tion to L-threenine, the antibiotic contains other hydroxy amino acids: D-allothreenine, D-serine, trans-3-hydroxyproline, and erythro- $\beta$ -hydroxyleucine.

The mild alkaline hydrolysis of the antibiotic (0.1 N NaOH, 37° C, 1 h, or 5% solution, 20° C, 0.5 h) gave a linear 11-membered polypeptide with C-terminal L-cis-3-hydroxyproline (acid of the antibiotic A-128-OP) which migrated to the anode on electrolysis in a buffer with pH 6.5.

On qualitative and quantitative amino-acid analysis of an acid hydrolysate of the acid of the antibiotic A-128-OP there was no L-threonine participating by its OH group in the formation of a lactone bond. Since it is known that some O-derivatives of serine undergo  $\beta$ elimination on treatment with alkaline agents with the formation of dehydroalanine [4-6], we assume that such elimination takes place in the alkaline treatment of the antibiotic A-128-OP. As a result of  $\beta$ -elimination, Lthreonine gives  $\beta$ -methyldehydroalanine, which decomposes in a similar manner to dehydroalanine [7, 8] on acid hydrolysis. The presence of  $\beta$ -methyldehydroalanine in the A-128-OP acid was shown by two methods.

1. The acid of the antibiotic was hydrogenated over a Pd/C catalyst, after which the  $\beta$ -methyldehydroalanine had been converted into  $\alpha$ -aminobutyric acid. The latter was identified in the hydrolysate of the reduced A-128-OP acid by thin-layer chromatography and electrophoresis and on a "Hitachi" type KLA-3B automatic amino acid analyzer (Fig. 1).

2. It is known that sodium sulfite adds quantitatively to the double bond of dehydroalanine forming cysteic acid [9]. The action of sodium sulfite on the acid of the antibiotic gave a substance with a higher mobility to the anode on electrophoresis in electrolytes with pH 2.4 and 6.5 than the initial A-128-OP acid. This

M. V. Lomonosov Moscow State University. Translated from Khimiya Prirodnykh Soedinenii, No. 4, pp. 544-546, July-August, 1971. Original article submitted March 24, 1971.

© 1973 Consultants Bureau, a division of Plenum Publishing Corporation, 227 West 17th Street, New York, N. Y. 10011. All rights reserved. This article cannot be reproduced for any purpose whatsoever without permission of the publisher. A copy of this article is available from the publisher for \$15.00.

shows the addition of a  $SO_3$  H grouping to the double bond of the dehydro amino acid. In addition to this, in an acid hydrolysate of the  $SO_3$  H derivative of the A-128-OP acid an amino acid was found with electrophoretic and chromatographic properties close to those of cysteic acid.

Thus, the results given show that  $\beta$ -methyldehydroalanine is formed in the mild alkaline treatment of the antibiotic peptidocyclolactone A-128-OP.

## LITERATURE CITED

- 1. A. B. Silaev, G. S. Katrukha, Zh. P. Trifonova, and I. G. Sinyavskaya, 2nd All-Union Biochemical Conference (abstracts) [in Russian], Tashkent (1969).
- 2. A. B. Silaev, G. S. Katrukha, Zh. P. Trifonova, I. G. Sinyavskaya, T. M. Melent'eva, M. Bakhra, and A. N. Polin, 7th International Symposium on the Chemistry of Natural Compounds (abstracts) [in Russian], Riga (1970).
- 3. A. B. Silaev, G. S. Katrukha, Zh. P. Trifonova, and T. M. Melent'eva, Khim. Prirodn. Soedin., 7, 130 (1971).
- 4. I. Photaki, J. Amer. Chem. Soc., <u>85</u>, 1123 (1963).
- 5. P. J. Hamrick and C. R. Hauser, J. Org. Chem., <u>26</u>, 4199 (1961).
- 6. D. H. Strumeyer, W. White, and D. E. Koshland, Jr., Proc. Natl. Acad. Sci. U.S., 50, 931 (1963).
- 7. A. Patchornik and M. Sokolovsky, J. Amer. Chem. Soc., <u>86</u>, 1206 (1964).
- 8. E. Gross and J. Morell, J. Amer. Chem. Soc., <u>89</u>, 2791 (1967).
- 9. H. Weiner, W. N. White, D. G. Hoare, and D. E. Koshland, Jr., J. Amer. Chem. Soc., 88, 3851 (1966).